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## Structure Reports

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# Tetraaquabis[5-(pyridin-4-yl)tetrazolido $N^5$ -oxide- $\kappa N^2$ ]manganese(II)

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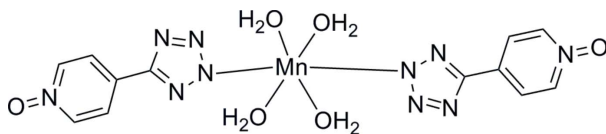
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(C-C) = 0.004$  Å;  $R$  factor = 0.037;  $wR$  factor = 0.080; data-to-parameter ratio = 10.6.

The title compound,  $[Mn(C_6H_4N_5O)_2(H_2O)_4]$ , is isotopic with its Zn, Ni and Cd analogues reported recently. In the crystal, the  $Mn^{II}$  cations are coordinated by four O atoms from four aqua ligands and two N atoms from two 5-(pyridin-4-yl)tetrazolido  $N^5$ -oxide ligands in a distorted octahedral coordination environment. The asymmetric unit consists of one  $Mn^{II}$  cation located on a crystallographic twofold axis, and two crystallographically independent water molecules and one  $N$ -donor ligand in general positions. The discrete complex molecules are arranged in alternating rows parallel to  $[100]$  and are linked by  $O-H\cdots N$  and  $O-H\cdots O$  hydrogen bonds into a three-dimensional network.

## Related literature

For related structures, see: Yang *et al.* (2009); Yu *et al.* (2004*a,b*). For the coordination properties of tetrazolate ligands, see: Aromí *et al.* (2011).



## Experimental

### Crystal data

 $[Mn(C_6H_4N_5O)_2(H_2O)_4]$   
 $M_r = 451.29$ 

 Monoclinic,  $C2/c$   
 $a = 21.828$  (2) Å

 $b = 7.0620$  (9) Å  
 $c = 11.3229$  (13) Å  
 $\beta = 96.515$  (10)°  
 $V = 1734.1$  (3) Å<sup>3</sup>  
 $Z = 4$ 
Mo  $K\alpha$  radiation $\mu = 0.82$  mm<sup>-1</sup> $T = 293$  K $0.32 \times 0.25 \times 0.20$  mm

### Data collection

 Bruker APEX area-detector  
 diffractometer  
 Absorption correction: multi-scan  
 (SADABS; Sheldrick, 1996)  
 $T_{\min} = 0.782$ ,  $T_{\max} = 0.849$ 

 5196 measured reflections  
 1530 independent reflections  
 1149 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.049$ 

### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.037$   
 $wR(F^2) = 0.080$   
 $S = 1.05$   
 1530 reflections  
 145 parameters

 H atoms treated by a mixture of  
 independent and constrained  
 refinement
 $\Delta\rho_{\text{max}} = 0.21$  e Å<sup>-3</sup> $\Delta\rho_{\text{min}} = -0.25$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$O2-H9\cdots O3^i$	0.87 (3)	1.80 (3)	2.658 (3)	170 (3)
$O1-H11\cdots O3^{ii}$	0.80 (3)	1.98 (3)	2.770 (3)	173 (3)
$O2-H10\cdots O3^{iii}$	0.87 (3)	1.88 (3)	2.751 (3)	172 (3)
$O1-H12\cdots N3^{iv}$	0.81 (3)	2.05 (3)	2.861 (3)	171 (3)

 Symmetry codes: (i)  $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$ ; (ii)  $x + \frac{1}{2}, y - \frac{1}{2}, z$ ; (iii)  $-x + \frac{1}{2}, y + \frac{1}{2}, -z + \frac{1}{2}$ ; (iv)  $x, -y, z - \frac{1}{2}$ .

Data collection: APEX2 (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXL97.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NC2281).

## References

- Aromí, G. L., Barrios, A., Roubeau, O. & Gamez, P. (2011). *Coord. Chem. Rev.* **255**, 485–546.  
 Bruker (2004). APEX2 and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.  
 Yang, W. B., Lin, X., Blake, A. J., Wilson, C., Hubberstey, P., Champness, N. R. & Schröder, M. (2009). *CrystEngComm*, **11**, 67–81.  
 Yu, Z.-X., Wang, X.-P. & Feng, Y. (2004a). *Acta Cryst.* **C60**, m194–m196.  
 Yu, Z. X., Wang, X. P., Feng, Y. Y. & Zhong, X. H. (2004b). *Inorg. Chem. Commun.* **7**, 492–494.

## supplementary materials

*Acta Cryst.* (2012). E68, m1130 [doi:10.1107/S1600536812032618]

**Tetraaquabis[5-(pyridin-4-yl)tetrazolido  $N^5$ -oxide- $\kappa N^2$ ]manganese(II)**

**Xiang Jing and Ya Luo**

**Comment**

Ligands based on tetrazolates have attracted wide attention because of their versatile coordination modes and therefore, a large number of metal complexes containing these types of ligands are reported in literature (Aromí *et al.*, 2011). Several of them have been prepared *via in situ* synthesis of nitriles and azides. In view of this we have reacted 4-cyanopyridine 1-oxide with manganese chloride and sodium azide, which results in the formation of crystals of the title compound that is isotopic to its Zn, Ni and Cd analogs (Yu *et al.*, 2004*a,b*; Yang *et al.*, 2009).

The coordination geometry of the Mn atom is a slightly distorted octahedron formed by the coordination of four water molecules and two tetrazolate ligands (Fig. 1). The oxygen atoms from the four water molecules form a square planar arrangement around the Mn center and the tetrazolate ligands coordinate *via* the N atom to the Mn cations. The discrete complex molecules are arranged in alternating rows parallel to [100] and are linked by O—H $\cdots$ N and O—H $\cdots$ O hydrogen bonds into a three-dimensional network. (Fig. 2 and Table 1)

**Experimental**

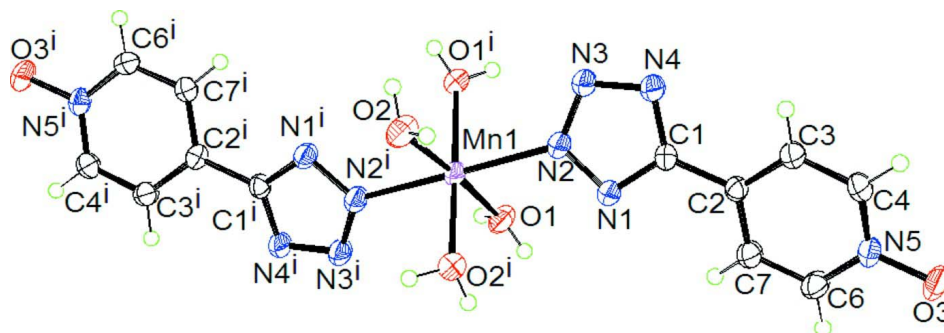
The mixture of 4-cyanopyridine 1-oxide (0.42 mmol, 50.3 mg),  $MnCl_2 \cdot 6H_2O$  (0.50 mmol, 116.5 mg) and  $NaN_3$  (0.70 mmol, 45.6 mg) in 15 ml EtOH and  $H_2O$  (v/v = 2:1) were heated in a 25 ml bomb at 393 K for 3 d, then cooled to room temperature at a rate of 6 K  $h^{-1}$ . Colorless block-shaped crystals suitable for X-ray analysis were obtained in a yield of 40% based on the ligand 4-cyanopyridine 1-oxide. The product was washed with EtOH and  $H_2O$ , and then air-dried at ambient temperature. Elemental analysis for  $C_{12}H_{16}N_{10}O_6Mn$  found: C 31.75, H 3.52, N 30.97; calculated: C 31.94, H 3.57, N 31.04. Selected IR (KBr,  $cm^{-1}$ ): 3118 (*w*), 3058(*w*), 1532 (*m*), 1462 (*m*), 1439 (*m*), 1215 (*s*), 1193 (*s*), 856 (*s*).

**Refinement**

Hydrogen atoms were placed in calculated positions (C—H 0.93 Å;  $U = 1.2U_{eq}C$ ), and were included in the refinement in the riding model approximation. The hydrogen atoms of aqua ligands were located and refined.

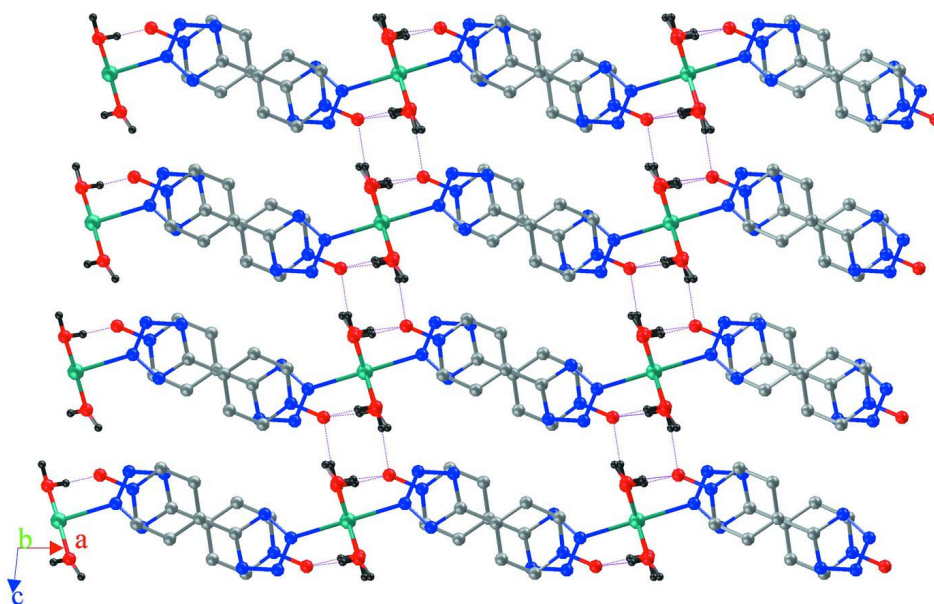
**Computing details**

Data collection: *APEX2* (Bruker, 2004); cell refinement: *SAINT* (Bruker, 2004); data reduction: *SAINT* (Bruker, 2004); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97* (Sheldrick, 2008).



**Figure 1**

ORTEP drawing of the title compound with labeling. Displacement ellipsoids are drawn at 30% probability level and H atoms are drawn as spheres of arbitrary radius. [Symmetry code:  $-x + 1, y, -z + 1/2$ ]



**Figure 2**

The packing diagram of the title compound with intermolecular H bonding along *ac* plane shown as dashed lines.

**Tetraaquabis[5-(pyridin-4-yl)tetrazolido *N*<sup>5</sup>-oxide- $\kappa$ N<sup>2</sup>]manganese(II)**

*Crystal data*

$[\text{Mn}(\text{C}_6\text{H}_4\text{N}_5\text{O})_2(\text{H}_2\text{O})_4]$

$M_r = 451.29$

Monoclinic,  $C2/c$

Hall symbol:  $-C\ 2yc$

$a = 21.828\ (2)\ \text{\AA}$

$b = 7.0620\ (9)\ \text{\AA}$

$c = 11.3229\ (13)\ \text{\AA}$

$\beta = 96.515\ (10)^\circ$

$V = 1734.1\ (3)\ \text{\AA}^3$

$Z = 4$

$F(000) = 924$

$D_x = 1.729\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 65 reflections

$\theta = 2.2\text{--}26.0^\circ$

$\mu = 0.82\ \text{mm}^{-1}$

$T = 293\ \text{K}$

Block, colourless

$0.32 \times 0.25 \times 0.20\ \text{mm}$

### Data collection

Bruker APEX area-detector diffractometer	5196 measured reflections
Radiation source: fine-focus sealed tube	1530 independent reflections
Graphite monochromator	1149 reflections with $I > 2\sigma(I)$
$\varphi$ and $\omega$ scans	$R_{\text{int}} = 0.049$
Absorption correction: multi-scan ( <i>SADABS</i> ; Sheldrick, 1996)	$\theta_{\text{max}} = 25.0^\circ$ , $\theta_{\text{min}} = 3.0^\circ$
$T_{\text{min}} = 0.782$ , $T_{\text{max}} = 0.849$	$h = -24 \rightarrow 25$
	$k = -8 \rightarrow 7$
	$l = -13 \rightarrow 13$

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.037$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.080$	$w = 1/[\sigma^2(F_o^2) + (0.0291P)^2]$
$S = 1.05$	where $P = (F_o^2 + 2F_c^2)/3$
1530 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
145 parameters	$\Delta\rho_{\text{max}} = 0.21 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -0.25 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Mn1	0.5000	0.16002 (8)	0.2500	0.0261 (2)
N1	0.35497 (9)	0.1653 (3)	0.2280 (2)	0.0300 (6)
N2	0.40483 (10)	0.1527 (3)	0.3078 (2)	0.0296 (5)
N3	0.38821 (10)	0.1262 (3)	0.4150 (2)	0.0328 (6)
N4	0.32672 (10)	0.1218 (3)	0.4086 (2)	0.0320 (6)
O1	0.47398 (9)	−0.0641 (3)	0.12179 (19)	0.0344 (5)
O2	0.52626 (10)	0.3978 (3)	0.3678 (2)	0.0380 (6)
O3	0.06246 (8)	0.1620 (3)	0.09684 (18)	0.0383 (5)
C1	0.30785 (11)	0.1464 (4)	0.2926 (2)	0.0236 (6)
C2	0.24297 (11)	0.1479 (3)	0.2420 (2)	0.0229 (6)
C3	0.19542 (12)	0.1790 (3)	0.3111 (3)	0.0296 (6)
H3	0.2044	0.1985	0.3924	0.036*
C4	0.13523 (12)	0.1815 (4)	0.2608 (3)	0.0311 (7)
H4	0.1037	0.2006	0.3084	0.037*
N5	0.12168 (9)	0.1565 (3)	0.1437 (2)	0.0282 (5)
C6	0.16617 (12)	0.1269 (4)	0.0737 (3)	0.0327 (7)

H6	0.1558	0.1112	−0.0077	0.039*
C7	0.22684 (12)	0.1197 (4)	0.1205 (3)	0.0304 (7)
H7	0.2573	0.0959	0.0713	0.036*
H9	0.5404 (16)	0.367 (4)	0.440 (3)	0.058 (5)*
H11	0.5008 (16)	−0.137 (4)	0.111 (3)	0.058 (5)*
H10	0.4956 (15)	0.475 (4)	0.376 (3)	0.058 (5)*
H12	0.4499 (16)	−0.069 (4)	0.061 (3)	0.058 (5)*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Mn1	0.0187 (3)	0.0362 (4)	0.0234 (4)	0.000	0.0021 (2)	0.000
N1	0.0206 (13)	0.0446 (14)	0.0245 (13)	−0.0004 (10)	0.0011 (10)	0.0032 (12)
N2	0.0205 (12)	0.0402 (13)	0.0275 (14)	0.0003 (10)	0.0003 (10)	0.0011 (12)
N3	0.0210 (13)	0.0499 (15)	0.0274 (14)	0.0011 (10)	0.0023 (11)	0.0025 (12)
N4	0.0210 (13)	0.0494 (15)	0.0258 (14)	0.0031 (10)	0.0038 (10)	0.0009 (12)
O1	0.0257 (13)	0.0483 (13)	0.0278 (12)	0.0033 (9)	−0.0032 (9)	−0.0070 (11)
O2	0.0309 (13)	0.0486 (14)	0.0334 (13)	0.0065 (9)	−0.0014 (10)	−0.0086 (11)
O3	0.0202 (10)	0.0496 (12)	0.0429 (13)	−0.0007 (9)	−0.0065 (9)	−0.0007 (11)
C1	0.0189 (14)	0.0254 (14)	0.0267 (16)	−0.0009 (11)	0.0033 (12)	−0.0010 (13)
C2	0.0205 (14)	0.0222 (14)	0.0257 (15)	−0.0017 (11)	0.0011 (12)	0.0036 (14)
C3	0.0236 (15)	0.0409 (16)	0.0245 (15)	−0.0020 (12)	0.0033 (12)	−0.0017 (14)
C4	0.0232 (15)	0.0427 (18)	0.0286 (16)	0.0007 (12)	0.0086 (12)	−0.0010 (15)
N5	0.0188 (12)	0.0314 (12)	0.0339 (14)	−0.0004 (10)	0.0011 (10)	0.0002 (12)
C6	0.0275 (16)	0.0464 (18)	0.0235 (16)	0.0020 (13)	−0.0005 (13)	−0.0061 (14)
C7	0.0229 (15)	0.0418 (17)	0.0268 (16)	0.0040 (12)	0.0045 (13)	−0.0013 (14)

*Geometric parameters (Å, °)*

Mn1—O1	2.179 (2)	O2—H10	0.87 (3)
Mn1—O1 <sup>i</sup>	2.179 (2)	O3—N5	1.341 (3)
Mn1—O2	2.180 (2)	C1—C2	1.467 (3)
Mn1—O2 <sup>i</sup>	2.180 (2)	C2—C3	1.387 (4)
Mn1—N2 <sup>i</sup>	2.248 (2)	C2—C7	1.394 (4)
Mn1—N2	2.248 (2)	C3—C4	1.371 (4)
N1—C1	1.335 (3)	C3—H3	0.9300
N1—N2	1.336 (3)	C4—N5	1.338 (3)
N2—N3	1.319 (3)	C4—H4	0.9300
N3—N4	1.336 (3)	N5—C6	1.338 (4)
N4—C1	1.342 (3)	C6—C7	1.371 (3)
O1—H11	0.80 (3)	C6—H6	0.9300
O1—H12	0.81 (3)	C7—H7	0.9300
O2—H9	0.87 (3)		
O1—Mn1—O1 <sup>i</sup>	86.82 (11)	Mn1—O2—H9	115 (2)
O1—Mn1—O2	175.99 (9)	Mn1—O2—H10	113 (2)
O1 <sup>i</sup> —Mn1—O2	96.97 (8)	H9—O2—H10	104 (3)
O1—Mn1—O2 <sup>i</sup>	96.97 (8)	N1—C1—N4	112.3 (2)
O1 <sup>i</sup> —Mn1—O2 <sup>i</sup>	175.99 (9)	N1—C1—C2	123.7 (2)
O2—Mn1—O2 <sup>i</sup>	79.27 (12)	N4—C1—C2	124.0 (2)

O1—Mn1—N2 <sup>i</sup>	88.23 (8)	C3—C2—C7	117.2 (2)
O1 <sup>i</sup> —Mn1—N2 <sup>i</sup>	89.86 (8)	C3—C2—C1	122.1 (2)
O2—Mn1—N2 <sup>i</sup>	90.47 (8)	C7—C2—C1	120.6 (2)
O2 <sup>i</sup> —Mn1—N2 <sup>i</sup>	91.55 (8)	C4—C3—C2	120.8 (3)
O1—Mn1—N2	89.86 (8)	C4—C3—H3	119.6
O1 <sup>i</sup> —Mn1—N2	88.23 (8)	C2—C3—H3	119.6
O2—Mn1—N2	91.55 (8)	N5—C4—C3	120.2 (3)
O2 <sup>i</sup> —Mn1—N2	90.47 (8)	N5—C4—H4	119.9
N2 <sup>i</sup> —Mn1—N2	177.37 (12)	C3—C4—H4	119.9
C1—N1—N2	104.0 (2)	C4—N5—C6	121.0 (2)
N3—N2—N1	110.08 (19)	C4—N5—O3	118.9 (2)
N3—N2—Mn1	129.10 (16)	C6—N5—O3	120.1 (2)
N1—N2—Mn1	120.71 (17)	N5—C6—C7	120.7 (3)
N2—N3—N4	109.4 (2)	N5—C6—H6	119.6
N3—N4—C1	104.2 (2)	C7—C6—H6	119.6
Mn1—O1—H11	115 (3)	C6—C7—C2	120.1 (3)
Mn1—O1—H12	133 (2)	C6—C7—H7	120.0
H11—O1—H12	105 (4)	C2—C7—H7	120.0
C1—N1—N2—N3	0.4 (3)	N3—N4—C1—N1	0.1 (3)
C1—N1—N2—Mn1	176.92 (17)	N3—N4—C1—C2	178.7 (2)
O1—Mn1—N2—N3	123.9 (2)	N1—C1—C2—C3	−162.3 (3)
O1 <sup>i</sup> —Mn1—N2—N3	37.1 (2)	N4—C1—C2—C3	19.3 (4)
O2—Mn1—N2—N3	−59.8 (2)	N1—C1—C2—C7	17.0 (4)
O2 <sup>i</sup> —Mn1—N2—N3	−139.1 (2)	N4—C1—C2—C7	−161.5 (3)
N2 <sup>i</sup> —Mn1—N2—N3	80.5 (2)	C7—C2—C3—C4	0.0 (4)
O1—Mn1—N2—N1	−51.82 (19)	C1—C2—C3—C4	179.3 (2)
O1 <sup>i</sup> —Mn1—N2—N1	−138.64 (19)	C2—C3—C4—N5	−1.0 (4)
O2—Mn1—N2—N1	124.43 (19)	C3—C4—N5—C6	0.6 (4)
O2 <sup>i</sup> —Mn1—N2—N1	45.15 (19)	C3—C4—N5—O3	−179.1 (2)
N2 <sup>i</sup> —Mn1—N2—N1	−95.22 (18)	C4—N5—C6—C7	0.7 (4)
N1—N2—N3—N4	−0.4 (3)	O3—N5—C6—C7	−179.6 (2)
Mn1—N2—N3—N4	−176.50 (17)	N5—C6—C7—C2	−1.6 (4)
N2—N3—N4—C1	0.2 (3)	C3—C2—C7—C6	1.2 (4)
N2—N1—C1—N4	−0.3 (3)	C1—C2—C7—C6	−178.1 (2)
N2—N1—C1—C2	−178.9 (2)		

Symmetry code: (i)  $-x+1, y, -z+1/2$ .

### Hydrogen-bond geometry ( $\text{\AA}$ , $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O2—H9 $\cdots$ O3 <sup>ii</sup>	0.87 (3)	1.80 (3)	2.658 (3)	170 (3)
O1—H11 $\cdots$ O3 <sup>iii</sup>	0.80 (3)	1.98 (3)	2.770 (3)	173 (3)
O2—H10 $\cdots$ O3 <sup>iv</sup>	0.87 (3)	1.88 (3)	2.751 (3)	172 (3)
O1—H12 $\cdots$ N3 <sup>v</sup>	0.81 (3)	2.05 (3)	2.861 (3)	171 (3)

Symmetry codes: (ii)  $x+1/2, -y+1/2, z+1/2$ ; (iii)  $x+1/2, y-1/2, z$ ; (iv)  $-x+1/2, y+1/2, -z+1/2$ ; (v)  $x, -y, z-1/2$ .